Preparation and Characterization of Activated Carbon from Palm Tree Leaves Impregnated with Zinc Chloride for the Removal of Lead (II) from Aqueous Solutions

Mutasim H. Elhussien¹,* , Rashida M. Hussein², Sumia A. Nimir¹, Mawia H. Elsaim³

¹Department of Chemistry, Faculty of Education, Nile Valley University, Atbara, Sudan
²Department of Chemistry, Inaya Medical College, Riyadh, Kingdom of Saudi Arabia
³Department of Chemistry, Faculty of Science and Technology, Abdulatif Alhamed University of Technology, Merowe, Sudan

Email address: mutasim1966@yahoo.com (M. H. Elhussien)

*Corresponding author

To cite this article:

Received: June 27, 2017; Accepted: July 10, 2017; Published: July 31, 2017

Abstract: Removal of Lead (II) from aqueous solution onto Activated carbon derived from Palm tree leaves (Phoenix dactylifera) as a novel precursor by chemical activation with ZnCl₂, at specific condition of carbonization temperature at 500°C and 1 hour as a holding time and detonated as AC(PTL)ZnCl₂. The obtained material was submitted to measurements of active surface area, Fourier Transform Infrared Spectroscopy (FTIR) analysis and scanning electron microscope (SEM). Batch-adsorption studied had been carried out to examine the adsorption capacity of the AC(PTL)ZnCl₂ for the removal of Lead (II) from aqueous solution. The effect of various process parameters like pH, initial metal concentration, adsorbent dose, contact time, and temperature on the efficiency of Pb (II) removal was investigated. Maximum adsorption of Pb (II) on AC(PTL)ZnCl₂ (86.4%) was observed at pH 5.81. The optimum conditions for adsorbent dose and temperature were determined as 0.030 g and 25°C, respectively. Initial Pb (II) concentrations has important effect on AC(PTL)ZnCl₂ in the studied range (25–125 mg/L) where the removal percentage increases as the metal ion decrease. The highest percentage removal of concentration corresponding to the maximum adsorption was found to be 82.75 ± 3.65. The adsorption equilibrium data was well explained by Langmuir, Freundlich and Dubinin-Radushkevich isotherm. The parameters suggested that the adsorption of Pb (II) on prepared AC(PTL)ZnCl₂ is physical adsorption. The Langmuir isotherm which fitted best for the experimental data obtained showed (Correlation Coefficient, R² = 0.9463) higher than the (Freundlich isotherm, R² = 0.7554), but more close to (Dubinin-Radushkevich isotherm, R² = 0.9260). The porous characteristics and adsorption efficiencies of prepared AC(PTL)ZnCl₂ were also investigated.

Keywords: Adsorption Isotherms, Activated Carbon, Palm Tree Leaves, Lead (II), Surface Area, Pore Size, SEM, FTIR

1. Introduction
The lack of clean drinking water and rise in environmental water pollution are a growing global concern. So availability of safe drinking water is still distance dream to many around the world, especially in developing countries. Increasing of human activities and industrialization have led to a wide range of physical, chemical and biological pollutants entering water, bodies and affecting human lives, thousands of chemicals have been identified in drinking water supplies around the world hazardous to human health.

Nowadays water pollution by heavy metals is fast growing due to natural processes and increasing human activities which include mining, agriculture, and manufacturing industries [1]. There for, must be work to create a healthier, safer, and cleaner environment for everyone by helping to guarantee clean drinking water, preserving raw water sources and ensuring the optimal operation water purification processes. Heavy metals which
are natural components of the Earth's crust are usually associated with toxicity. Exposure to heavy metals, even at trace level, is known to be a risk for human beings [2]. There are 59 elements classified as heavy metals and out of these five are considered to be highly toxic and hazardous heavy metals [3]. These are lead (Pb) cadmium (Cd), chromium (Cr), copper (Cu), and zinc (Zn), have a potentially damaging effect on human physiology. Metal toxicity can result from drinking-water contamination (e.g. lead pipes), increased ambient air concentrations near sources of emission, or ingestion via the food chain [4]. The increased use of heavy metals in industry has resulted in increased availability of metallic substances in natural water sources [5, 6].

Lead, Pb (II) is among those contaminants that must be removed from water, due to its high toxicity and tendency to accumulate in tissues of living organisms [7]. Lead can occur in different forms such as organic and inorganic forms. Lead is the most significant toxin of the trace metal ions, and human exposure to the inorganic forms is through ingestion of food and water. As the result, there has been a great deal of attention given to new technologies for removal of heavy metal ions from contaminated water. Many technologies like adsorption, precipitation, membrane filtration, and ion-exchange have been used to remove metal pollutants from water [8]. However, adsorption has proven to be economical and efficient for removing heavy metals, organic pollutants and dyes from polluted waters [9].

Numerous researchers have extensively studied the preparation of activated carbon from various types of precursors and evaluate the performance of such developed materials by adsorbing of different heavy metals [10]. A wide variety of activated carbons have been prepared from agricultural waste such as coconut shells, cotton stalk, sugarcane bagasse, coir pith, straw [11-13] and also from non-renewable materials like coal, wood, peat [14], and many others.

The present study is mainly focused on preparing the micro-porous activated carbon from palm tree leaves (as novel material) by chemical activation method using Zinc Chloride (ZnCl₂) as a chemical agent. Also, evaluated the removal efficiency of Pb (II) ions from the aqueous phase using the prepared activated carbon.

2. Materials and Methods

The carbonaceous precursor used for preparation of activated carbon (AC) was Palm tree leaves, collected from Alsalam Park in Riyadh, Kingdom of Saudi Arabia. Prior to use, sample was washed gently with water to remove impurities present on the surface and then dried for one week. All chemical reagents used in this work are procured from Pharmchem Haryana and General Drug House (p) Ltd Daryagani, New Delhi, (India).

2.1. Preparation of Activated Carbon

Palm tree leave, collected, dried, crushed and grinded in a ball mill. The grinded sample was sieved to obtain the particles of uniform size, 1.0– 1.5 mm. The precursor obtained was washed to remove surface bounded impurities and dried at 105°C for 24 h. Raw material of about 100 g was impregnated with ZnCl₂ for 24 h in impregnation ratio of 1: 1.5 at 50°C to achieve well penetration of the chemical into the interior of the precursor.

About 10 g of the Palm tree leave impregnated with the activating agent ZnCl₂ and the impregnation was transferred to a stain less steel reactor (150 mm length and 40 mm diameter) with narrow ports at both ends. The tube is placed inside an electric horizontal tubular furnace (Bysakh & Co) controlled by the proportional integral derivative (PID) controller. The furnace tube dimensions are 800 mm length and 50 mm in diameter. The temperature (500°C) of the reaction zone was measured by the chromel – alumel (K-type) thermocouple. The pyrolysis of the samples was carried out under continuous flow of N₂ gas (200 ml/min). The carbonized samples were cooled to room temperature under inert atmosphere and washed with water followed by 0.1 M hydrochloric acid (HCl) to remove the residual chemical agents until the pH value of the rinsed water was neutral. The adsorbents prepared was denoted as AC(PTL)ZnCl₂ throughout the work.

![Figure 1. Images of activated carbon obtained from palm tree leave after carbonization process.](image)

2.2. Evaluation of Activated Carbon

The activated carbon sample was evaluated three ways:

2.2.1. Specific Surface Area Measured by Liquid Nitrogen Adsorption

Nitrogen (N₂) gas adsorption-desorption isotherms on prepared AC(PTL)ZnCl₂ at liquid nitrogen temperature (–195.6°C) were carried out using an automatic adsorption unit, Autosorb – 1(Quanta chrome). The samples were degassed at 200°C for 3h prior to analysis so as to remove any adsorbed moisture or other impurities bounded to the surface of the sample. Surface area (ST) values were calculated from the experimental adsorption isotherm over a relative pressure range of 0.01 to 0.3 using the standard BET (Brunauer, Emmett and Teller) method. The accuracy of
measurements performed by this equipment was ± 5%. The BET equation is given as [15].

\[
P/V(P^0 - P) = \left(1/V_mC\right) + (C-1) P/P_o V_mC
\]  

where, \( V \) is the volume adsorbed at STP (cm\(^3\)g\(^{-1}\)), \( V_m \) is the volume of monolayer capacity at STP(cm\(^3\)g\(^{-1}\)), and the term \( C \), the BET constant, is related to the energy of adsorption in the first adsorbed layer and its value is an indication of the magnitude of the adsorbent-adsorbate interactions. The total pore volume (\( V_t \)) of the samples was determined as the total volume of N\(_2\) gas adsorbed at high relative pressure (0.995) [16].

2.2.2. Fourier Transform Infrared Spectroscopy

The surface functional groups of the ACs were estimated by Fourier Transform Infrared (FTIR) spectroscopy (Nicolet 6700, Thermo Scientific, USA). FTIR spectra of the activated and inactivated AC(PTL)ZnCl\(_2\) samples were recorded within 400 – 4000 cm\(^{-1}\). The transmission spectra of the samples were recorded using the KBr pellet. About 1.0 – 2.0% of the sample was mixed with dry KBr and ground in mortar. Then the samples were transferred to hydraulic presses. The pellets which are homogenous and transparent in appearance were dried overnight at 100°C, and then inserted into the IR sample holder for the analysis.

2.2.3. Scanning Electron Microscopy (SEM)

SEM analysis of samples was performed by using (JSM-6380LA) scanning electron microscope. The SEM instrument was operated at 5 kV/SE (Accelerating Voltage Machine), and 50°C inclination. Before analysis, all samples were ground and coated in a sputter coating unit (Edwards Vacuum Components Ltd., Sussex, England) to reduce charging and improve the secondary electron signals for imaging. The micrographs were recorded using photographic techniques.

2.3. Effect of Process Parameters

Batch adsorption experiments for adsorption of Pb (II) on AC(PTL)ZnCl\(_2\) were conducted using aqueous solutions of the metal. A stock solution of 1000 mg/L of Pb (II) was prepared by dissolving appropriate weight of Pb (NO\(_3\))\(_2\) (General Drug House)(p)Ltd. Daryaganj, New Delhi (India) in 1% (v/v) nitric acid solution in a 1.0 L volumetric flask. The stock solution was diluted as required to obtain different concentrations of Pb (II) solutions. For each run, a definite amount of AC(PTL)ZnCl\(_2\) was added to 25 ml of Pb (II) solution taken in 100 ml Erlenmeyer flasks. All the adsorption experiments were carried out at constant temperature of 25°C.

2.3.1. Effect of pH

The effect of pH on metal adsorption was monitored over a pH range of 1.5 to 9, using electronic pH - Meter (pH 211). In this study, 25 ml of separate solutions 25 mg/L Pb (II) was transferred into 100 ml conical flasks agitated at 100 rpm for 180 min with 0.03 g AC(PTL)ZnCl\(_2\) at 25. The mixture was filtered and the filtrate analyzed for residual metal ions using FAAS (model 932AB).

2.3.2. Effect of Contact Time

The investigation was done to determine the optimum time for adsorption of lead ions using AC(PTL)ZnCl\(_2\). The biosorbent (0.03 g) was contacted with 25 ml separate solutions of 25 mg/L Pb (II) ions adjusted to pH 5.81. The solution was agitated at 100 rpm for different contact times (1-1440 min) at 25°C. The mixtures were filtered and residual metal ions in filtrate analyzed using FAAS.

2.3.3. Effect of Adsorbent Dosage

The effect of quantity of AC(PTL)ZnCl\(_2\) used was investigated by equilibrating 25 ml solutions of 25 mg/L Pb (II) ions adjusted to pH 5.81, with different masses of the adsorbent (0.01-0.05 g) agitated at 100 rpm for 180 min at 25. The mixture was filtered and the residual metal ions in filtrate analyzed using FAAS.

2.3.4. Effect of Initial Metal Concentration

Separate 25 ml solution of lead ions with different initial concentrations (25-125mg/L) were contacted with optimized adsorbent dosage 0.03 g pH 5.81. The mixtures were agitated at 100 rpm for 180 min at 25°C. The mixtures were filtered and filtrate analyzed for residual metal ions.

2.3.5. Effect of Temperature

The effect of temperature of 25 ml solutions, 25 mg/L Pb (II) ions on adsorption was monitored over a range of 25 to 55°C. The investigations were carried out at optimized adsorbent dose, contact time, pH, and agitation time. The mixtures were filtered and filtrate analyzed for residual lead ions using FAAS.

2.4. Equilibrium Adsorption Studies

The adsorption isotherms for the adsorption of Pb (II) onto AC(PTL)ZnCl\(_2\) were investigated by varying the initial Pb (II) concentration from 25 to 125 mg/L at optimized adsorbent dose, contact time, \( pH \), temperature, and stirring speed (established after optimization of working parameters). The equilibrium data were investigations using Langmuir, Freundlich, Dubinin – Radushkevich Isotherms models.

The data obtained in the batch modes were used to calculate the equilibrium metal ion adsorptive amounts (mg/g) were calculated using the following expressions [17, 18]:

\[
q_e = \left(\frac{V}{w}\right) \left(C_o - C_e\right)
\]

\[
\% \text{Removal} = 100 \left(\frac{C_o - C_e}{C_o}\right)
\]

Where \( q_e \) is the amount of the heavy metal ion adsorbed by the AC(PTL)ZnCl\(_2\) sample in (mg/g), \( C_o \) and \( C_e \) are the initial and equilibrium concentrations of the metal ion respectively in (mg/L), \( V \) is the volume of the solution treated in (L) and \( w \) is the adsorbent mass of (AC(PTL)ZnCl\(_2\)) in (g).
3. Results and Discussion

3.1. Evaluation of Activated Carbon

3.1.1. Specific Surface Area Measured by Liquid Nitrogen Adsorption

Surface area and pore properties of the prepared AC(PTL)ZnCl$_2$ are tabulated in Table 1. Figure (2 & 3) represents the incremental mesopore volume and Micropore size distribution for the prepared AC(PTL)ZnCl$_2$ respectively. BET surface plot for N$_2$ gas adsorption on the activated prepared sample is shown in Figure 4.

<table>
<thead>
<tr>
<th>Physical properties of the prepared AC(PTL)ZnCl$_2$</th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>BET Surface Area</td>
<td>299.842± 5.0495% m$^2$/g</td>
</tr>
<tr>
<td>Langmuir Surface Area</td>
<td>370.8677m$^2$/g</td>
</tr>
<tr>
<td>$q_m$ (at STP)</td>
<td>68.8786 m$^3$/g</td>
</tr>
<tr>
<td>Pore volume</td>
<td>0.129132 cm$^3$/g</td>
</tr>
<tr>
<td>Pore Size (Adsorption average pore width)</td>
<td>17.2267 Å</td>
</tr>
<tr>
<td>Correlation Coefficient, (R)</td>
<td>0.999585</td>
</tr>
</tbody>
</table>

**Figure 2.** Incremental mesopore volume for prepared AC(PTL)ZnCl$_2$

**Figure 3.** Micropore size distribution for AC(PTL)ZnCl$_2$
3.1.2. Fourier Transform Infrared Spectroscopy (FTIR)

The FTIR analysis of prepared AC(PTL)ZnCl$_2$ reveals a simple spectrum, shown on figure 5. The broad peaks at around 3423.43 cm$^{-1}$ is assigned to presence of OH groups (adsorbed moisture) on surface. Silane (Si – H) groups in prepared AC(PTL)ZnCl$_2$ were confirmed by the peak at around 2368.66 and 2368.74 cm$^{-1}$ respectively. A little shift of the bands corresponding to silane groups was observed in prepared ACs. The presence of amine groups are confirmed by the peaks observed at 1612.39 and 1570.40 cm$^{-1}$ while the band at 1108.13 cm$^{-1}$ is due to the C – N stretching.

3.1.3. Scanning Electron Microscope (SEM)

The prepared activated carbon AC(PTL)ZnCl$_2$ was examined by Scanning Electron Microscope (SEM) to analyze the surface of the adsorbents. SEM micrographs of the chemically activated carbons by ZnCl$_2$, was presented in Figure 6 respectively. Well-developed porous surface was observed at higher magnification. The pores observed from SEM images are having diameter in micrometer ($\mu$m) range.
These pores are considered as channels to the microporous network. From the figure, it can be observed that all the adsorbents have rough texture with heterogeneous surface and a variety of randomly distributed pore size.

3.2. Effect of Process Parameters

3.2.1. Effect of pH

Solution pH is one of the important process parameters that significantly influences the adsorption of Pb (II) on adsorbent [19, 20]. AC(PTL)ZnCl$_2$ was employed for Pb (II) removal at different pH values (1.5 – 9.0). The preliminary experiments were performed at definite experimental conditions (initial Pb (II) concentration – 25.0 mg/L, adsorbent dose – 0.030g, contact time – 3.0 h, and temperature – 25°C). Figure 7 shows the higher removal of Pb (II) at 5.81 pH.

It is evident from the figure that the percentage adsorption is higher at acidic pH (5.5 – 5.81), reaching maximum of about 86.4% ± 4.2. But it decreased gradually with the decrease in pH. The maximum adsorption capacities observed at pH 5.81.0 (18.00mg/g). Adsorption values at pH greater than 6.9 gives rise to lead precipitation, so that active sites would have achieved saturation above a certain range of pH values. The results obtained are in close agreement with reported studies [19, 21].

These obtained results could be explained as, in acidic media i.e. when the pH of solution less than 4, the elevated hydrogen ions (H$^+$) will rival the Pb (II) ions from reaching the free available sites on the adsorbent surface and prevent the Pb (II) ions from binding with activated carbon surface due to the repulsive forces. Besides that, it is also noted when the pH values reaches 5.81, the adsorption process enhanced and removal percentage get maximum quantities regarding the lower hydrogen ions (H$^+$) exist, and on the sense the Pb (II) ions will get better chance to occupy the free sites of the active surface. Later on, when the pH precede toward basic region and elevated above 5.81, the recorded values of removal ions increase significantly specially at 7 and 9 pH to record higher values. This is attributed to Pb (II) ions precipitated regarding to the formation of lead hydroxide anions due to the dissolution of Pb(OH)$_2$; as a result [18, 21].

3.2.2. Effect of Adsorbent Dose

The effect of adsorbent dose on removal percentage of Pb (II) using AC(PTL)ZnCl$_2$ was illustrated in Figure 8. Different doses of adsorbents ranging from 0.01 – 0.050 g were considered and other process parameters were maintained constant (pH – 5.81Pb (II) concentration – 25.0 mg/L, Agitation speed – 100 rpm, contact time – 3.0 h, and temperature – 25°C).

An increase in adsorption capacity with increasing adsorbent dose up to a maximum of 0.03 g giving the corresponding optimum percentage removal of 78.48%. On the other hand, it is found that any further addition over the above mentioned weight (0.03g) will not make any enhancement in the adsorption process, where almost negligible increase of removal efficiency over the specific adsorbent dose. The initial increase in adsorption capacity with increasing adsorbent mass is explained by the increase in the number of exchangeable sites for metal ion adsorption, after which equilibration was attained [18, 22].
3.2.3. Effect of Contact Time

The effect of contact time on the removal of lead (II) is shown in Figure 9. The amount of the adsorbed metal ions increases with increasing time until it levels off after some (180 min). A constant adsorption is indicative of equilibration due to saturation of adsorption sites. Rapid adsorption of metal ions during the initial stages was due to the large initial concentration gradient between the adsorbate in solution and the number of available vacant sites on the adsorbent surface [18, 19, 23].

3.2.4. Effect of Initial Lead (II) Concentration

The initial concentration of metal ions in the solution is an important parameter since the metal ions concentration change over a broad range in effluents applications.

The batch adsorption experiments were carried out with different initial concentrations \( C_i \) (25, 50, 75, 100, and 125). The variation of percentage removal of Pb (II) ions with different initial concentration for the prepared activated carbon AC(PTL)ZnCl₂ was clarified in Figure 10. The figure shows a excellent performance of the prepared activated carbon at equilibrium state and clarify the optimum Pb (II) ions initial concentration used at confined experimental conditions.

It is also evidently observed that the percentage removal of the Pb (II) ions is sufficiently high, (86.4%) at low concentration (25 mg/L) and gradually decreases as the concentration increases [18]. This could be attributed to the fact that after the formation of mono-ionic layer at low concentration over the activated carbon surface, further formation of the layer is highly hindered at higher concentration and is due to interaction between Pb (II) ions on the surface and in the solution [18, 23].

3.2.5. Effect of Temperature

Effect of temperature on Pb (II) removal percentage on AC(PTL)ZnCl₂ was shown in Figure 11. Experiments were performed in the temperature range 25 – 55°C at constant Pb (II) concentration (25.0 mg/L), adsorbent dose (0.030 g), contact time (3.0 h) and pH (5.81) for AC(PTL)ZnCl₂. Adsorption of Pb (II) on the adsorbent witnessed the increase in removal percentage and adsorption capacity \( q_e \) at 25°C, while in the range from 35 to 55°C, no significant change with further increase of temperature. The figure represents high rate of adsorption at at 25. However low
rate of Pb (II) adsorption was observed at 55°C within the studied range of temperatures. The adsorption of Pb (II) on AC(PTL)ZnCl₂ decreased from 83.3 to 11.0%, with the rise in temperature from 25 to 55. This indicates that the removal of Pb (II) onto AC(PTL)ZnCl₂ is exothermic [18, 19, 24].

The high removal percentage of Pb (II) by AC(PTL)ZnCl₂ may be due to the high diffusion rate of metal ions into the pores as the surface area and pore volume of the adsorbent were large. However, at high temperature the kinetic energy of Pb (II) species is low and hence contact between the metal ions and the active sites of AC is in sufficient, resulting in reduced removal efficiency [25, 26, 27].

3.3. Adsorption Isotherms

The successful representation of the dynamic adsorptive separation of solute from solution by an adsorbent depends upon a good description of the equilibrium between the two phases. Adsorption equilibrium is established when the amount of solute being adsorbed onto the adsorbent is equal to the amount being desorbed [28]. The equilibrium adsorption isotherms were depicted by plotting solid phase concentration (qₑ) against liquid phase concentration (Cₑ) of solute.

Adsorption isotherm explains the interaction between adsorbate and adsorbent and is critical for design of adsorption process. The Langmuir, Freundlich, and D-R isotherms are the most frequently used models to describe the experimental data of adsorption. In the present work these four isotherms were applied to investigate the adsorption process of Pb (II) on prepared AC(PTL)ZnCl₂ at different conditions of process parameters.

3.3.1. Langmuir Isotherm

The Langmuir isotherm is applicable to homogeneous sorption where the sorption of each solute molecule on to the surface has equal sorption activation energy and is represented as follows [29, 30]:

\[ qₑ = \frac{K_L Cₑ}{1 + a_0 Cₑ} \]  \hspace{1cm} (4)

Where qₑ is the solid phase sorbate concentration at equilibrium, K_L and a₀ are the Langmuir isotherm constants.

The linear form of Langmuir equation is given as:

\[ \frac{Cₑ}{qₑ} = \frac{1}{K_L a_0} + \left(\frac{1}{a_0 K_L}\right) Cₑ \]  \hspace{1cm} (5)

The adsorption data were analyzed according to the linear form of equation (equation 5). The plots of \( Cₑ/qₑ \) versus \( Cₑ \) are linear which indicate that the adsorption data fitted reasonably to the Langmuir isotherm (Figure 12). The constants were evaluated from the slope \( a_0/K_L \) and intercept \( 1/K_L \), where \( K_L/a_0 \) gives the theoretical monolayer saturation capacity \( Q_w \). The Langmuir constants obtained are summarized in Table 2.

The positive value of \( a_0 \) (0.166), obtained for AC(PTL)ZnCl₂ indicates the efficiency of Langmuir model to explain the adsorption process. The adsorption data for AC(PTL)ZnCl₂ was well fitted to the Langmuir equation with reasonably high regression coefficient (0.9463).

3.3.2. Freundlich Isotherm

The most important multilayer adsorption isotherm for heterogeneous surfaces is the Freundlich isotherm which is characterized by the heterogeneity factor \( 1/n \), and is represented by the equation:

\[ qₑ = K_F Cₑ^{1/n} \]  \hspace{1cm} (6)

where, \( qₑ \) is the solid phase concentration in equilibrium, \( Cₑ \) is the liquid phase sorbent concentration at equilibrium, \( K_F \) is the Freundlich constant and \( 1/n \) is the heterogeneity factor. The magnitude of \( n \) gives an indication on the favorability of adsorption. It is generally stated that the values of \( n \) in the range 2 – 10 represent good, 1 – 2 moderately good, and less than 1 poor adsorption characteristics [31]. The Freundlich isotherm is an empirical equation based on an exponential distribution of adsorption sites and energies. The linear form of Freundlich equation is:

\[ \ln qₑ = \ln K_F + \left(\frac{1}{n}\right) \ln Cₑ \]  \hspace{1cm} (7)

where, the intercept \( \ln K_F \) is a measure of adsorption capacity, and the slope \( 1/n \) is the adsorption intensity. The values of \( K_F \) and \( n \) were calculated from the intercept and slope of the plots \( \ln qₑ \) against \( \ln Cₑ \) (Figure 13). The Freundlich isotherm describes reversible adsorption and was
not restricted to the formation of monolayer. The Freundlich isotherm showed a poor fit to the adsorption data than Langmuir isotherm, with regression coefficient of 0.7554 compared to 0.9463 in Langmuir. This indicates that Pb (II) adsorption on AC(PTL)ZnCl₂ was monolayer adsorption. Freundlich sorption isotherm constants were determined and are summarized in Table 2.

3.3.3. Dubinin – Radushkevich Isotherm

The linear form of Dubinin – Radushkevich (DR) isotherm equation is presented as [32]:

\[
\ln q_e = \ln q_0 - 2B_0 RT \ln \left(1 + 1/C_s\right) \quad (8)
\]

A plot of \(\ln q_e\) against \(RT \ln (1 + 1/C_s)\) in Figure 14 showed straight line and indicates a good fit of the isotherm to the experimental data. The apparent energy \((E)\) of adsorption from DR isotherm model can be estimated using the equation given below:

\[
E = 1/ \left(2B_0\right)^{1/2} \quad (9)
\]

The DR isotherm constants and mean free energy were presented in Table 2. It was found that the DR isotherm model gives a satisfactory fit for Pb (II) adsorption data for AC(PTL)ZnCl₂ with good regression coefficient (0.926). The magnitude of the activation energy can be used to determine the type of adsorption process. It easy to have a general insight into the undergoing process. If the value of \(E \leq 8\) kJ/mol, the adsorption type can be explained by physisorption, and if \(E\) is in the range of \(8 – 18\) kJ/mol then the adsorption type is ion exchange [33], and chemical-reaction-controlled when \(E > 50\) kJ/mol \([23, 34]\). The low value of apparent energies obtained \((0.09414\) kJ/mol\) confirms that the adsorption of Pb (II) on prepared AC(PTL)ZnCl₂ is physical adsorption [35].

![Figure 14. Dubinin – Radushkevich isotherms of AC(PTL)ZnCl₂ for the Pb (II) removal at 25°C.](image)

### Table 2. Langmuir, Freundlich, and Dubinin – Radushkevich isotherm constants for the Pb (II) removal on the prepared AC(PTL)ZnCl₂ at 25°C.

<table>
<thead>
<tr>
<th>Parameters</th>
<th>Langmuir</th>
<th>Freundlich</th>
<th>Dubinin Radushkevich</th>
</tr>
</thead>
<tbody>
<tr>
<td>(K_L) (L/mg)</td>
<td>9.843</td>
<td>0.166</td>
<td>94.14</td>
</tr>
<tr>
<td>(a_L) (L/mol)</td>
<td>0.988</td>
<td>0.7554</td>
<td>0.926</td>
</tr>
<tr>
<td>(Q_L) (mg/g)</td>
<td>59.30</td>
<td>0.9463</td>
<td>2</td>
</tr>
<tr>
<td>(1/n)</td>
<td>0.351</td>
<td>3.9</td>
<td></td>
</tr>
<tr>
<td>(B_0) X10⁻²</td>
<td>0.166</td>
<td>0.141</td>
<td></td>
</tr>
<tr>
<td>(E) (J/mol)</td>
<td>94.14</td>
<td>0.926</td>
<td></td>
</tr>
</tbody>
</table>

4. Conclusion

The main conclusions which are achieved from this research are:

1. The prepared activated carbon, AC(PTL)ZnCl₂, used in this study is found to have a high capability for removing the Pb (II) ions from aqueous solution.
2. The adsorption method is successfully applied to remove Pb (II) ions from aqueous solution. The removal efficiency of lead (II) ions was found to be dependent on various parameters such as pH of the solution, initial lead concentration, mass of adsorbent and contact time.
3. The results obtained showed that the removal efficiency of the Pb (II) ions by the prepared activated carbon, AC(PTL)ZnCl₂, increased with time, which is significant in the first stage with an instantaneous removal of Pb (II) ions, after that the curves become as plateau after 180 minutes.
4. The experimental results showed that the removal efficiency of the Pb (II) ions by AC(PTL)ZnCl₂ activated carbon decreases with increase in initial Pb (II) ions concentration, and increases with increase in mass of adsorbent.
5. The pH of the solution played an important role in influencing the ability of adsorbents towards the metal ion. The increase in the pH of solution from 2 to 5.81 led to increase in the removal efficiency.
6. The optimum conditions which have been obtained in adsorption process for treating the aqueous phase were 50ppm for initial lead concentration, 5.81 pH for solution acidity, and 15 g/L for mass of adsorbent.
7. The equilibrium data obtained for the prepared activated carbon, AC(PTL)ZnCl₂, fitted to the Langmuir and Dubinin – Radushkevich models more than the Freundlich model.

Acknowledgements

We would like to thank Dr. Mohammed Elfaki at Chemistry Department, Faculty of Science, King Saud University, for supporting and assistance to conduct this study.

References


